Draw the reaction energy diagram, and label the products, reactants, transition states, and intermediates.
1.


- 3-Step reaction
- $2^{\text {nd }}$ step is Rate-Determining-Step
- $3^{\text {rd }}$ step is faster than the $1^{\text {st }}$
- $1^{\text {st }}$ step is endothermic
- $2^{\text {nd }}$ step is exothermic
- $3^{\text {rd }}$ step is exothermic
- The overall reaction is exothermic
- 3-Step reaction
- $3^{\text {rd }}$ step is Rate-Determining-Step
- $1^{\text {st }}$ step is faster than the $2^{\text {nd }}$
- $1^{\text {st }}$ and $2^{\text {nd }}$ step are endothermic
- $3^{\text {rd }}$ step is exothermic
- The overall reaction is neutral

Reaction progress
3.


## Reaction progress

4. $\qquad$

## Reaction progress

5. $\square$

## Reaction progress

Draw mirror images of the following.
6.

8.

7.

9.


- 4-Step reaction
- $3^{\text {rd }}$ step is Rate-Determining Step
- $1^{\text {st }}$ step is endothermic
- $1^{\text {st }}$ and $2^{\text {nd }}$ step proceed at the same speed
- $3^{\text {rd }}$ step is exothermic
- $4^{\text {th }}$ step is exothermic
- The overall reaction is exothermic
- 3-Step reaction
- $3^{\text {rd }}$ step is Rate-Determining-Step
- $1^{\text {st }}$ step is faster than the $2^{\text {nd }}$
- $1^{\text {st }}$ and $3^{\text {rd }}$ steps are exothermic
- $2^{\text {nd }}$ step is endothermic
- The overall reaction is endothermic

10. 



14.

16.

18.


11.

13.

15.

17.


Label the chiral carbons with an asterisk (*), and determine whether the compound itself is chiral or achiral.
1.

4.

7. $\mathrm{H}_{3}$ C゙


8.


11.
3.

6.

9.

12.



13.
14.



19.

17. Br

20. $\mathrm{CH}_{3} \mathrm{CH}_{3} / / \prime, \mathrm{C}=\mathrm{C}=\mathrm{C}_{\mathrm{H}}^{\mathrm{Cl}}$




18.

